

# Reagents for Pharma Industry

## Chapter 2



**PanReac**   
**AppliChem**  
ITW Reagents



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## About Us

### *The Origin*

ITW Illinois Tool Works Inc. (NYSE: ITW) is a global industry company that delivers specialized expertise, innovative thinking and value-added products to meet critical customer needs in a variety of industries.

ITW, with approximately 14 billion dollars in global revenues, operates 7 major segments with businesses in 58 countries that employ approximately 50,000 employees. The company has a broad portfolio of more than 17,000 global patents and patent applications.

### *The ITW Reagents Division*

In 2010, the ITW Reagents division was born integrated by the companies Panreac Química SLU (Spain) and Nova Chimica Srl (Italy), and later on by AppliChem GmbH (Germany). The division offers the highest quality and innovative products for analysis, research and production applications.

ITW Reagents markets its products worldwide through an extensive distribution network to more than 80 countries under the PanReac AppliChem brand. It has two production plants in Darmstadt (Germany) and Barcelona (Spain).



### *We are Everywhere*

We can say that almost all products subject to human manipulation have undergone chemical analysis that guarantees their physical and chemical properties. Food, agrifood, medicines, cosmetics... and so many other products are subjected to chemical analysis. Our reagents can be found in any quality control and research laboratory.



#### **Our range of Laboratory Chemicals include:**

Analytical reagents  
Reagents for volumetric analysis  
Reagents and solvents for general applications  
Reagents and solvents for HPLC  
Reagents and solvents for GC  
Reagents for metallic traces analysis  
Analytical standards  
Reagents and solvents for specific applications  
Products for clinical diagnosis  
Products for microbiology

#### **Our range of Laboratory Biochemicals cover:**

Cell Biology / Cell Culture  
Protein Biochemistry and Electrophoresis  
Nucleic Acid Biochemistry  
General Biochemicals and Biological Buffers  
Special Biochemicals

## *Service & Benefits*

**Exceptional know-how** and a wide range of chemicals and biochemicals for a great diversity of applications.

**European production** committed to corporate social responsibility (CSR).

**Efficient global distribution network** to export our products worldwide to more than 80 countries.

**Qualified management team** fully committed to our business project.

## *Excellence*

Our products are strictly controlled in our laboratories and meet the highest quality requirements. A multi-site Integrated Management System for Quality, Environment and Safety is implemented in all activities and processes.

ISO 9001:2015

ISO 14001:2015

OHSAS 18001:2007



# Reagents for Pharma Industry

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## Introduction

The **Pharmaceutical Industry** discovers, develops, produces, and markets drugs or **pharmaceutical drugs** for use as medications.

Pharmaceutical companies may deal in **generic** or **brand medications** and medical devices.

They are subject to a variety of **laws** and **regulations** that govern the patenting, testing, safety, efficacy and marketing of drugs.

The pharmaceutical industry is largely driven by **scientific discovery** and **development**, in conjunction with **toxicological** and **clinical experience**.



Major differences exist between **large organizations** which engage in a broad range of drug discovery and development, manufacturing and quality control, marketing and sales and **smaller organizations** which focus on a specific aspect.



Most multinational pharmaceutical companies are involved in all these activities; however, they may specialize in one aspect based upon local market factors. Academic, public and private organizations perform scientific **research to discover and develop new drugs**. The biotechnology industry is becoming a major contributor to innovative pharmaceutical research. Often, collaborative agreements between research organizations and large pharmaceutical companies are formed to explore the potential of new drug substances.

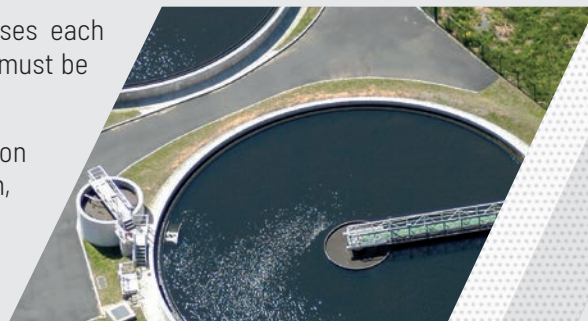
**Active drug substances** (APIs, Active Principle Ingredient) and **inert materials** (Excipients) are combined **during pharmaceutical manufacturing** to produce dosage forms of medicinal products (e.g. tablets, capsules, liquids, powders, creams and ointments). Drugs may be categorized by their manufacturing process and therapeutic benefits.





The different pharmaceutical manufacturing processes each have their own **environmental issues** and the wastes must be treated and controlled. **For example:**

- During **fermentation process**, the spent fermentation broth contains sugars, starches, proteins, nitrogen, phosphates and other nutrients with high biochemical oxygen demand (BOD), chemical oxygen demand (COD) and total suspended solids (TSS) with pH values ranging from 4 to 8.
- Also, wastes from **chemical synthesis** are complex due to the variety of hazardous materials, reactions and unit operations. These waste waters are high in BOD, COD and TSS, with varying acidity or alkalinity and pH values ranging from 1 to 11.



The **analysis laboratories** play a fundamental role in the pharmaceutical industries. **They are key pieces in:**



- Discovery and improvement of a **drug**.
- Development and optimization of **manufacturing processes**.
- **Quality control** of raw materials, intermediates and finished products.
- Quality control of **wastes**.

!

Depending on the type of analysis in which they are involved, **different types of laboratories** can be distinguished within the same pharmaceutical company. Besides, the **type of analysis** and the techniques used may be different (as shown on the next page).

In any case, the methods of analysis must be strictly validated and follow the requirements set by the **Pharmacopoeias** (Ph. Eur., USP, etc.) both in the analysis protocols and in the quality of the reagents to be used.

Our **portfolio** includes a wide range of products such as solvents, acids, bases and salts indicated for general analytical applications that **fulfil the requirements indicated in the Pharmacopoeias** (Ph. Eur. or USP) for the reagents to be used for analytical purposes.



## Types of Laboratories versus Methods of Analysis

Facility		R&D Centre		Manufacturing Plant Quality Control			Wastewater Plant
Laboratory		New molecules / Improvements of existing products	Analytical development	Raw Material (excipients & APIs)	In-process (intermediate product)	Final product	Water quality control
Methods of analysis	Chapter						
Amino acid analysis	5			●	●	●	
Ammonium	6/7						●
Approximate pH of solutions	1		●	●	●	●	●
Assay: Protein (Kjeldahl)	6		●	●			
Assay: Titration	6		●	●			
Assay: Water (KF)	6		●	●	●	●	
Atomic Absorption spectroscopy	2		●	●			
Biological assays	3		●	●			
Biological tests	3			●		●	
Clarity and opalescence of liquids	1		●	●		●	
Chlorinated compounds	7						●
Conductivity	1		●	●			
Degree of Coloration of Liquids	1		●	●			
Detergents (Surfactants)	7						●
Dissolution Test	1					●	
Electrophoresis	5	●	●	●	●	●	
Gas Chromatography	4	●	●	●		●	
ICP	2		●	●			
Identification	6		●	●		●	
IR	2	●	●	●		●	
Limit tests	6		●	●			
Liquid Chromatography	4	●	●	●		●	
Molecular mass distribution in dextrans	5			●	●		
Organic compounds (COD, DBO5, TOC)	7						●
Peptide identification by NMR spectrometry	5	●	●	●	●	●	
Peptide mapping	5	●		●	●	●	
Phosphates	6/7						●
Potentiometric determination of pH	1		●	●	●	●	●
Residual catalyzers (Metals, Cyanides)	7						●
Suspended matter	7						●
Thin Layer Chromatography	4	●	●	●			
UV	2	●	●	●			
Synthesis*	8	●					

\*not a method of analysis but reagents and solvents involved in synthesis procedures.

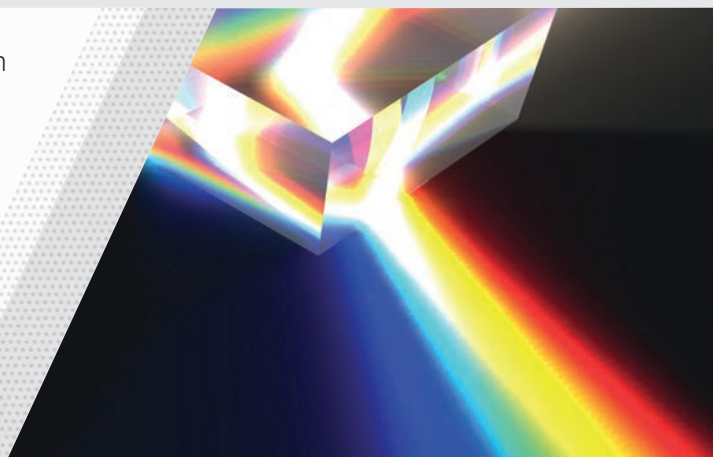
In the following sections we will describe the most common methods of analysis indicated in the pharmacopoeias and offer the most appropriate reagents for each method.



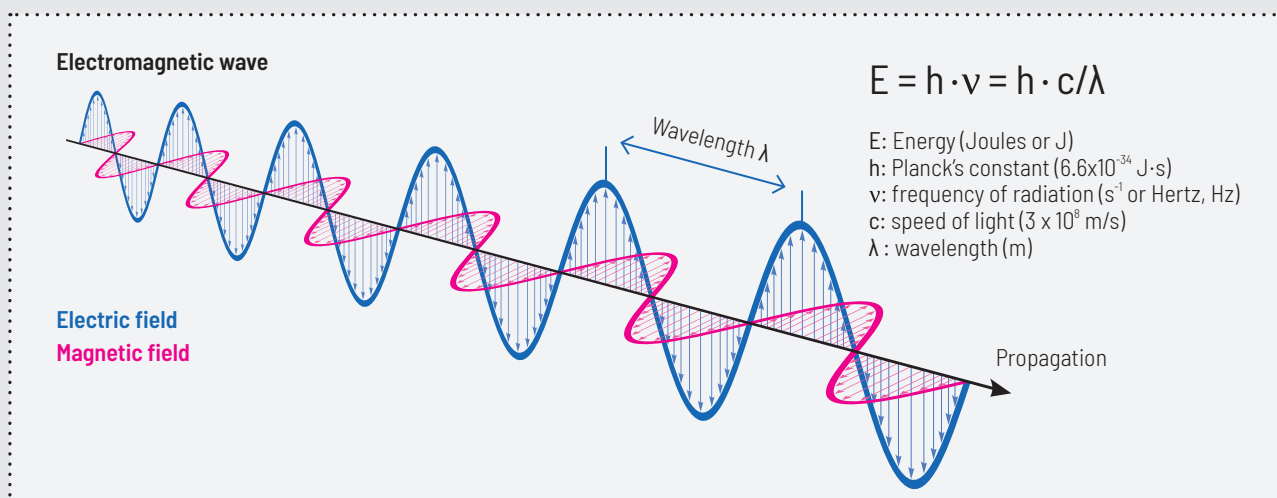


## Spectroscopy

**Spectroscopy** is the study of the interaction between matter and **electromagnetic radiation** involving either absorption, emission, or scattering of radiation by the material under study. It is often used in physical and analytical chemistry to obtain detailed information about the structure and chemical properties of substances through the spectrum emitted from or absorbed by them.



The electromagnetic radiation is composed of oscillating electric and magnetic fields that have the ability to transfer energy through space at high speed. The energy propagates as a wave.



**The electromagnetic spectrum** is the range of frequencies of electromagnetic radiation and their respective wavelengths and photon energies.

Electromagnetic radiation interacts with atoms and molecules. The nature of this interaction depends upon the energy of the electromagnetic radiation.

The spectroscopic techniques provide us with information about:

- The type of atom or molecule that is present
- How much of a particular atom or molecule is present
- The structure and bonding of the molecule

Spectroscopic techniques utilize the fact that:

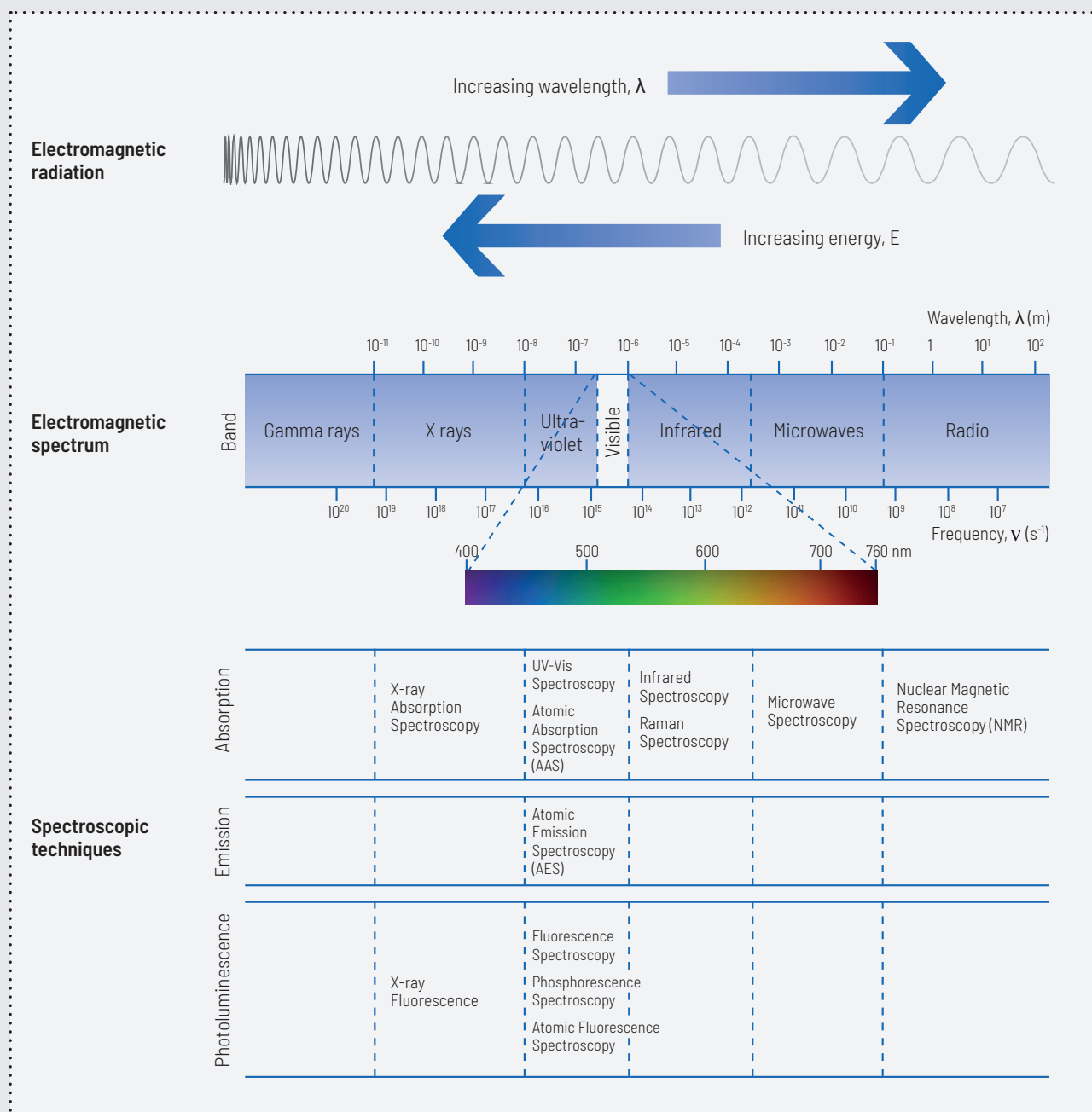
- Atoms or molecules absorb and emit electromagnetic radiation of specific energies
- Atoms and molecules undergo a change when they absorb electromagnetic radiation
- Different parts of the electromagnetic spectrum affect different parts of the atom or molecule

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The following image shows the type of spectroscopic technique depending on the type of energy transfer (absorption, emission or photoluminescence) and the range of electromagnetic spectrum used. The matter can be affected at different levels (molecule, atom or nucleus).



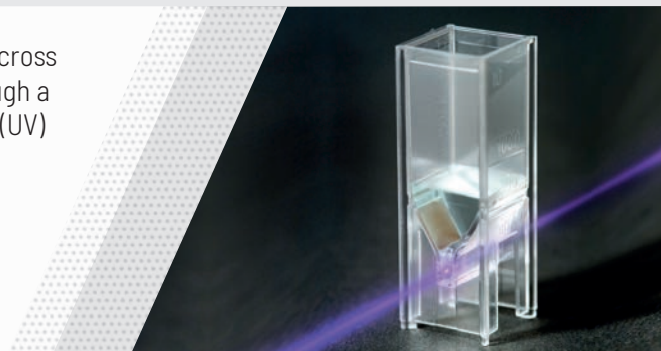
In this chapter we will talk about four types of spectroscopic techniques and also the recommended reagents and qualities that PanReac AppliChem can offer for each type of technique:

- UV-Vis spectroscopy
- Infrared spectroscopy
- Atomic Absorption spectroscopy
- Atomic Emission spectroscopy

## Ultraviolet-Visible Spectroscopy (UV-Vis)

This technique measures the **absorption of light** across the **ultraviolet and visible light wavelengths** through a **liquid sample**. Absorption of visible and ultraviolet (UV) radiation is associated with excitation of electrons, in both atoms and molecules, from lower to higher energy levels.

The absorption in the visible range directly affects the perceived **color** of the chemicals involved.



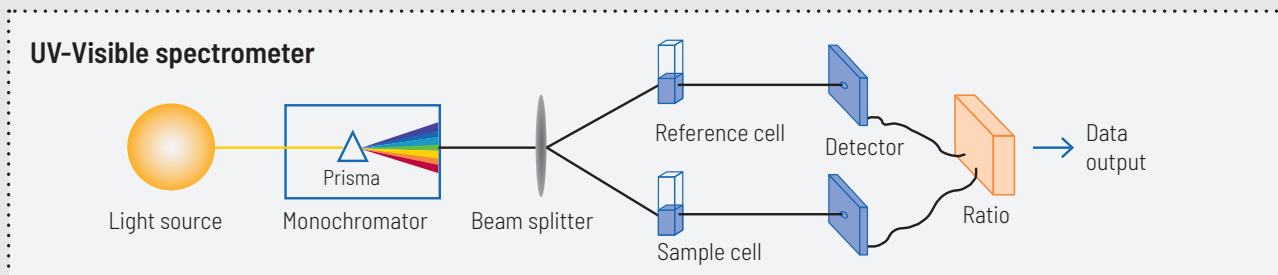
UV/Vis spectroscopy is routinely used in **analytical chemistry** for the **quantitative determination** of different analytes, such as transition metal ions, highly conjugated organic compounds, and biological macromolecules. Spectroscopic analysis is commonly carried out in solutions but solids and gases may also be studied.

A UV/Vis spectrophotometer may be used as a detector for **HPLC**. The presence of an analyte gives a response assumed to be proportional to the concentration.

### Apparatus

A **UV-visible spectrometer** can be used to measure the absorbance of ultraviolet or visible light by a sample, either at a single wavelength or perform a scan over a range in the spectrum. It consists of an optical system capable of producing monochromatic radiation in the range of 200–800 nm and a suitable device for measuring the absorbance. The UV region ranges **from 190 to 400 nm** and the visible region **from 400 to 800 nm**.

The technique can be used both quantitatively and qualitatively.



- The **light source** (a combination of tungsten/halogen and deuterium lamps) provides the visible and near ultraviolet radiation covering a range from 200 to 800 nm.
- The output from the light source is focused onto the **diffraction grating** or **monochromator** which splits the incoming light into its component colors of different wavelengths, like a prism but more efficiently.
- For liquids the sample is held in an optically flat, transparent container called a **cell** or **cuvette**. The **reference cell** or cuvette contains the solvent in which the sample is dissolved, commonly referred to as the blank. For each wavelength the intensity of light passing through both a reference cell ( $I_0$ ) and the sample cell ( $I$ ) is measured. If  $I$  is less than  $I_0$ , then the sample has absorbed some of the light. The absorbance ( $A$ ) of the sample is related to  $I$  and  $I_0$  according to the following equation:

$$A = \log_{10} \left( \frac{I_0}{I} \right)$$

- The **detector** converts the incoming light into a current; the higher the current, the greater the intensity. The chart recorder usually plots the absorbance against wavelength (nm) in the UV and visible section of the electromagnetic spectrum.





## UV-Visible Spectrum

The diagram in Figure 1 shows a simplified UV-visible absorption spectrum for 1,3-butadiene. Absorbance (on the vertical axis) is just a measure of the amount of light absorbed.

One can readily see what wavelengths of light are absorbed (peaks), and what wavelengths of light are transmitted (troughs). The higher the value, the more of a particular wavelength is being absorbed.

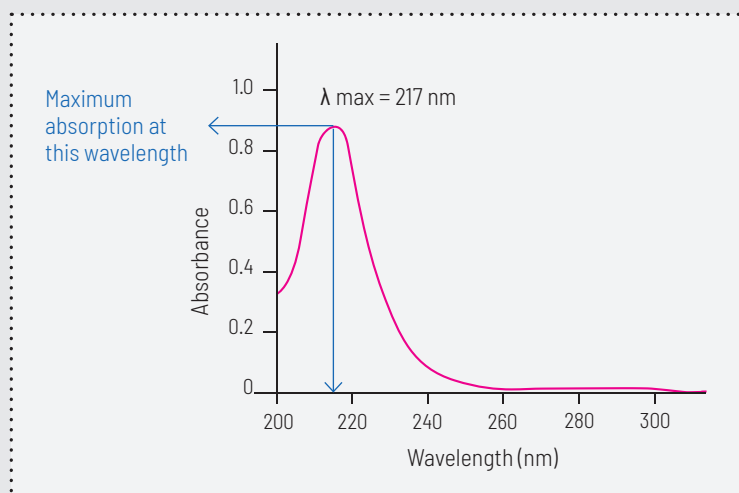


Figure 1. UV-Visible absorption spectrum for 1,3-butadiene

## The Beer-Lambert Law

The **Beer-Lambert** law states that the **absorbance** of a solution is directly proportional to the **concentration** of the absorbing species in the solution and the **path length**. Thus, for a fixed path length, UV/Vis spectroscopy can be used to determine the concentration of the absorber in a solution. It is necessary to know how quickly the absorbance changes with concentration. This can be taken from references (tables of molar extinction coefficients), or more accurately, determined from a calibration curve.

The Beer-Lambert Law can be expressed in the form of the following equation:

$$A = \epsilon c l$$

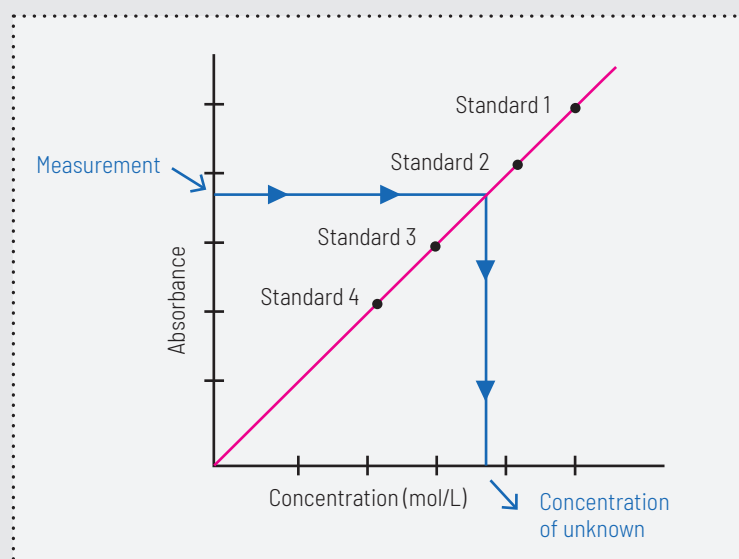
where

$A$  = absorbance (no units)

$\epsilon$  = molar extinction, which is constant for a particular substance at a particular wavelength ( $\text{dm}^3 \text{mol}^{-1} \text{cm}^{-1}$ )

$c$  = concentration of solution ( $\text{mol dm}^{-3}$ )

$l$  = optical path length, i.e. dimension of the cell or cuvette (cm)



If the absorbance of a series of sample solutions of known concentrations are measured and plotted against their corresponding concentrations, the plot of **absorbance versus concentration** should be **linear** if the Beer-Lambert Law is obeyed. This graph is known as a calibration graph.

A calibration graph can be used to determine the concentration of an unknown sample solution by measuring its absorbance, as in Figure 2.

Figure 2. Calibration graph

## UV/VIS Spectrophotometry Standards

UV/VIS spectrophotometers must be checked periodically to ensure their correct functioning, thereby guaranteeing the reliability of their results. The European Pharmacopoeia and the USP recommend certain solutions for the control of UV/VIS equipment that can be used to detect possible deviations arising from electrical, mechanical or optical problems.

The parameters to be checked in UV/VIS spectroscopy according to the European Pharmacopoeia and USP are as follows:

- Absorbance
- Accuracy of the wavelength
- Stray light limit
- Spectral resolution

PanReac AppliChem supplies four liquid standards that allow these parameters to be verified with the following advantages:

- Compliance with Ph. Eur. and USP
- Packaged in ready-to-use, disposable ampoules
- Certificate of analysis with each pack
- Three-year expiry date



Solutions for the control of:	Composition	Code	Package
<b>Absorbance</b>	Potassium Dichromate solution 4x10 mL standard ampoules (60.06 mg/L $K_2Cr_2O_7$ in $H_2SO_4$ 0.01N) 4x10 mL blank ampoules ( $H_2SO_4$ 0.01N)	395462.2590	8 x 10 ml
<b>Spectral resolution</b>	Toluene solution in n-Hexane 4x10 mL standard ampoules (0.02 mL toluene in 100 mL n-hexane) 4x10 mL blank ampoules (n-hexane)	395461.2590	8 x 10 ml
<b>Stray light</b>	Potassium Chloride solution 8x10 mL standard ampoules (12 g/L KCl in $H_2O$ )	395460.2590	8 x 10 ml
<b>Wavelength</b>	Holmium Oxide solution according to Ph. Eur. 8x10 mL Standard ampoules (40.0 g/L $Ho_2O_3$ in $HClO_4$ 10% v/v)	396070.2590	8 x 10 ml



## Solvents for UV-Vis

When preparing the solution of the substance to be examined, the solvent should completely solubilize the sample at the desired concentration. The solvent should be UV transparent at the measuring wavelength so as not to cause interference which could affect quantitative results.

When using volatile solvents stoppered cells should be employed to eliminate evaporation which can lead to changes in the solution concentration and ultimately affect quantitation.

PanReac AppliChem offers an excellent selection of UV-Vis transparent solvents to meet the most exigent spectrophotometric needs. We assure consistent, dependable performance and extremely low levels of optical impurities. Our solvents offer the following benefits:






































- More reliable sample identification, since solvents have highest chemical purity and contain no extraneous peaks
- Enhanced sensitivity, due to highest UV transmittance/lowest UV absorbance
- Reproducible absorption curve throughout the entire spectrum due to the highest batch-to-batch consistency

Product name	Assay (G.C.) min.	Non- volatile matter max.	Water max.	Fluorescence (as quinine) max.		UV transmittance (min.) at λ					Code	Pack	
				254 nm	365 nm								
Acetone	99.9 %	0.0003 %	0.2 %	-	2 ppb							361007.1611	☞ 1 L
						330 nm	335 nm	340 nm	345 nm	≥ 350 nm		361007.1612	☞ 2.5 L
						15%	60%	85%	95%	98%		361007.16153	☞ 4 L
												361007.0515	☞ 10 L
												361007.0537	☞ 30 L
Acetic Acid glacial	99.8 %	0.001 %	0.1 %	-	-	254 nm		280 nm		≥ 300 nm	361008.1611	☞ 1 L	
						25%		95%		98%		361008.1612	☞ 2.5 L
Acetonitrile	99.9 %	0.0004 %	0.02 %	1 ppb	1 ppb	193 nm	195 nm	200 nm		≥ 230 nm	361881.1611	☞ 1 L	
						55%	70%	90%	98%		361881.1612	☞ 2.5 L	
											361881.0314	☞ 2.5 L	
											361881.0516	☞ 25 L	
Benzene	99.8 %	0.0003 %	0.01 %	-	2 ppb	280 nm	285 nm	290 nm	300 nm	320 nm	≥ 340 nm	361192.1611	☞ 1 L
						25%	70%	80%	90%	95%	98%		
1-Butanol	99.9 %	0.0003 %	0.03 %	-	-	210 nm	220 nm	230 nm	240 nm	250 nm	≥ 270 nm	361082.1611	☞ 1 L
						25%	50%	75%	85%	94%	98%		
tert-Butyl Methyl Ether	99.9 %	0.0003 %	0.03 %	2 ppb	2 ppb	210 nm	235 nm	240 nm	255 nm	≥ 280 nm	363312.1611	☞ 1 L	
						20%	50%	60%	85%	98%		363312.1612	☞ 2.5 L
Carbon Disulfide	99.9 %	0.0005 %	0.005 %	-	-	390 nm	400 nm	410 nm		≥ 420 nm	361244.1611	☞ 1 L	
						50%	80%	90%		98%			
Chloroform stab. with ~ 150 ppm of amylene	99.8 %	0.0003 %	0.01 %	1 ppb	1 ppb	245 nm	255 nm	260 nm		≥ 280 nm	363101.1612	☞ 2.5 L	
						15%	70%	85%		98%			

>>



<<

Product name	Assay (G.C.) min.	Non- volatile matter max.	Water max.	Fluorescence (as quinine) max.		UV transmittance (min.) at $\lambda$						Code	Pack
				254 nm	365 nm								
<b>Chloroform stab. with ethanol</b>	99.0 %	0.0003 %	0.01 %	1 ppb	1 ppb	245 nm 20%	250 nm 50%	257 nm 80%	260 nm 85%	270 nm 90%	$\geq 280$ nm 98%	361252.1611	 1 L
<b>Cyclohexane</b>	99.9 %	0.0003 %	0.01 %	1 ppb	1 ppb	210 nm 15%	220 nm 50%	230 nm 80%	240 nm 90%	250 nm 98%	$\geq 260$ nm 99%	361250.1611	 1 L
												361250.1612	 2.5 L
												361250.0515	 10 L
												361250.0537	 30 L
<b>Dichloromethane stab. with ~ 20 ppm of amylene</b>	99.9 %	0.0003 %	0.01 %	1 ppb	1 ppb	235 nm 40%	240 nm 75%	245 nm 90%	255 nm 98%	$\geq 260$ nm 99%		361254.1611	 1 L
												361254.1612	 2.5 L
												361254.16153	 4 L
												361254.0516	 25 L
<b>Diethyl Ether stab. with ethanol</b>	98.0 %	0.0003 %	0.02 %	2 ppb	0.5 ppb	236 nm 50%	252 nm 80%	280 nm 94%	$\geq 300$ nm 99%			362551.1611	 1 L
												362551.0537	 30 L
<b>N,N-Dimethylacet- amide</b>	99.9 %	0.0005 %	0.03 %	-	-	280 nm 50%	290 nm 70%	310 nm 89%	320 nm 93%	$\geq 360$ nm 98%		363145.1611	 1 L
												363145.1612	 2.5 L
<b>N,N-Dimethylfor- mamide</b>	99.9 %	0.0003 %	0.05 %	-	-	275 nm 60%	290 nm 80%	300 nm 90%	$\geq 330$ nm 98%			361785.1611	 1 L
												361785.1612	 2.5 L
<b>Dimethyl Sulfoxide</b>	99.9 %	0.001 %	0.05 %	-	-	270 nm 30%	280 nm 63%	290 nm 70%	310 nm 80%	330 nm 94%	$\geq 350$ nm 98%	361954.1611	 1 L
												361954.1612	 2.5 L
<b>Ethanol absolute</b>	99.9 %	0.0003 %	0.1 %	2 ppb	1 ppb	210 nm 35%	220 nm 55%	235 nm 80%	245 nm 90%	$\geq 270$ nm 98%		361086.1611	 1 L
												361086.1612	 2.5 L
												361086.16153	 4 L
<b>Ethanol 96% v/v</b>	96 %	0.0005 %	-	-	-	210 nm 35%	220 nm 55%	230 nm 72%	240 nm 90%	$\geq 270$ nm 98%		361085.1611	 1 L
												361085.0537	 30 L
<b>Ethyl Acetate</b>	99.9 %	0.0003 %	0.01 %	2 ppb	1 ppb	255 nm 20%	257 nm 32%	260 nm 50%	263 nm 80%	265 nm 90%	$\geq 270$ nm 98%	361318.1611	 1 L
												361318.1612	 2.5 L
												361318.16153	 4 L
												361318.0515	 10 L
<b>n-Heptane</b>	99.0 %	0.0003 %	0.005 %	1 ppb	1 ppb	210 nm 50%	220 nm 80%	230 nm 92%	$\geq 245$ nm 98%			362062.1611	 1 L
												362062.1612	 2.5 L
												362062.0314	 5 L
												362062.0537	 30 L
<b>n-Hexane</b>	99.0 %	0.0003 %	0.005 %	2 ppb	2 ppb	200 nm 20%	210 nm 60%	220 nm 85%	225 nm 90%	230 nm 94%	$\geq 245$ nm 98%	362063.1611	 1 L
												362063.1612	 2.5 L
<b>n-Hexane 95%</b>	98.5 % (isomers)	0.0003 %	0.01 %	-	-	200 nm 20%	210 nm 60%	220 nm 80%	230 nm 94%	$\geq 245$ nm 98%		363242.1611	 1 L
												363242.1612	 2.5 L
												363242.0515	 10 L
												363242.0537	 30 L
<b>Hexane, alkanes mixture</b>	95.0 % (isomers)	0.0005 %	0.01 %	-	-	210 nm 40%		220 nm 85%		$\geq 254$ nm 99%		361347.1612	 2.5 L
























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# Reagents for Pharma Industry

## Chapter 2



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Product name	Assay (G.C.) min.	Non- volatile matter max.	Water max.	Fluorescence (as quinine) max.		UV transmittance (min.) at $\lambda$					Code	Pack	
				254 nm	365 nm								
Isooctane	99.5 %	0.0003 %	0.005 %	1 ppb	1 ppb	210 nm 50%	220 nm 80%	235 nm 90%	$\geq 245$ nm 98%	362064.1611	 1 L		
										362064.1612	 2.5 L		
										362064.16153	 4 L		
Methanol	99.9 %	0.0005 %	0.03 %	1 ppb	1 ppb	210 nm 30%	220 nm 55%	230 nm 75%	240 nm 90%	$\geq 260$ nm 98%	361091.1611	 1 L	
											361091.1612	 2.5 L	
											361091.16153	 4 L	
n-Pentane	99.5 %	0.0003 %	0.005 %	1 ppb	1 ppb	200 nm 40%	210 nm 70%	215 nm 80%	220 nm 90%	$\geq 240$ nm 98%	362006.1611	 1 L	
											362006.1612	 2.5 L	
1-Propanol	99.8 %	0.0002 %	0.1 %	-	-	210 nm 15%	230 nm 70%	240 nm 80%	250 nm 94%	$\geq 260$ nm 98%	361885.1611	 1 L	
											361885.1612	 2.5 L	
2-Propanol	99.9 %	0.0003 %	0.05 %	-	-	217 nm 50%	232 nm 80%	242 nm 90%	250 nm 95%	$\geq 260$ nm 98%	361090.1611	 1 L	
											361090.1612	 2.5 L	
											361090.16153	 4 L	
Propionitrile	99.9 %	0.0005 %	0.03 %	-	2 ppb	235 nm 94%		250 nm 97%		$\geq 290$ nm 99%	365732.1611	 1 L	
Tetrachloro-ethylene	99.9 %	0.0005 %	0.01 %	-	2 ppb	295 nm 50%	300 nm 80%	305 nm 85%	350 nm 89%	$\geq 400$ nm 94%	361455.1611	 1 L	
											361455.1612	 2.5 L	
Tetrahydrofuran	99.9 %	0.0002 %	0.02 %	-	-	240 nm 30%	245 nm 50%	260 nm 70%	265 nm 80%	275 nm 90%	$\geq 310$ nm 99%	361736.1611	 1 L
												361736.1612	 2.5 L
Toluene	99.9 %	0.0003 %	0.01 %	-	2 ppb	288 nm 32%	290 nm 50%	293 nm 63%	300 nm 80%	310 nm 90%	$\geq 350$ nm 98%	361745.1611	 1 L
												361745.1612	 2.5 L
												361745.16153	 4 L
Water	-	0.0003 %	-	1 ppb	0.5 ppb	200 nm 98%	210 nm 98%	254 nm 99%	$\geq 300$ nm 99%	361074.1611	 1 L		
										361074.1612	 2.5 L		

All solvents are filtered through 0.2  $\mu$ m and bottled under nitrogen atmosphere. APHA color:  $\leq 10$



## Solvents for UV-Vis in Returnable Containers

The consumption of solvents in 10 L and 30 L returnable stainless steel containers represents a fantastic opportunity to contribute to environmental sustainability.

It also involves a significant saving in the consumption of packaging materials as a result of the long-lasting and resistant nature of stainless steel containers, which therefore simplifies laboratory waste management.

Download our brochure "Solvents in Returnable Containers" and find out more at [https://www.itwreagents.com/download\\_file/brochures/A176/en/A176\\_en.pdf](https://www.itwreagents.com/download_file/brochures/A176/en/A176_en.pdf)





### Types of barrels & Withdrawal systems

#### 10 L – code BIN010TRF

Material	Stainless steel AISI 304
Total Capacity	12.2 L
Diameter	229 mm
Height	360 mm
Wall thickness	1 mm
Opening	Thread GL 45, internal opening 29 mm
Weight	2.4 kg
Admissible working pressure	max. 0.3 bar
UN Transport approval	1A1 / X 2.0 / 900

#### Accessories

Description	Code
Withdrawal system for manual pressurizing	DIS10RF



#### 30 L – code BIN030RF

Material	Stainless steel AISI 304
Total Capacity	33 L
Diameter	363 mm
Height	435 mm
Wall thickness	1.5 mm
Opening	Central bung 2" with Tri-Sure screw cap
Weight	10 kg
Admissible working pressure	max. 5 bar
UN Transport approval	1A1 / X 2.0 / 900

#### Accessories

Description	Code
Withdrawal system for manual pressurizing	DINOX3RF
Hand pump ball	BOMBIXRF
Withdrawal system for compressed air or nitrogen supply pressurizing	GOINOXRF
Metal wrench for drums with 2" and 3/4" closures Tri-Sure	100004RF













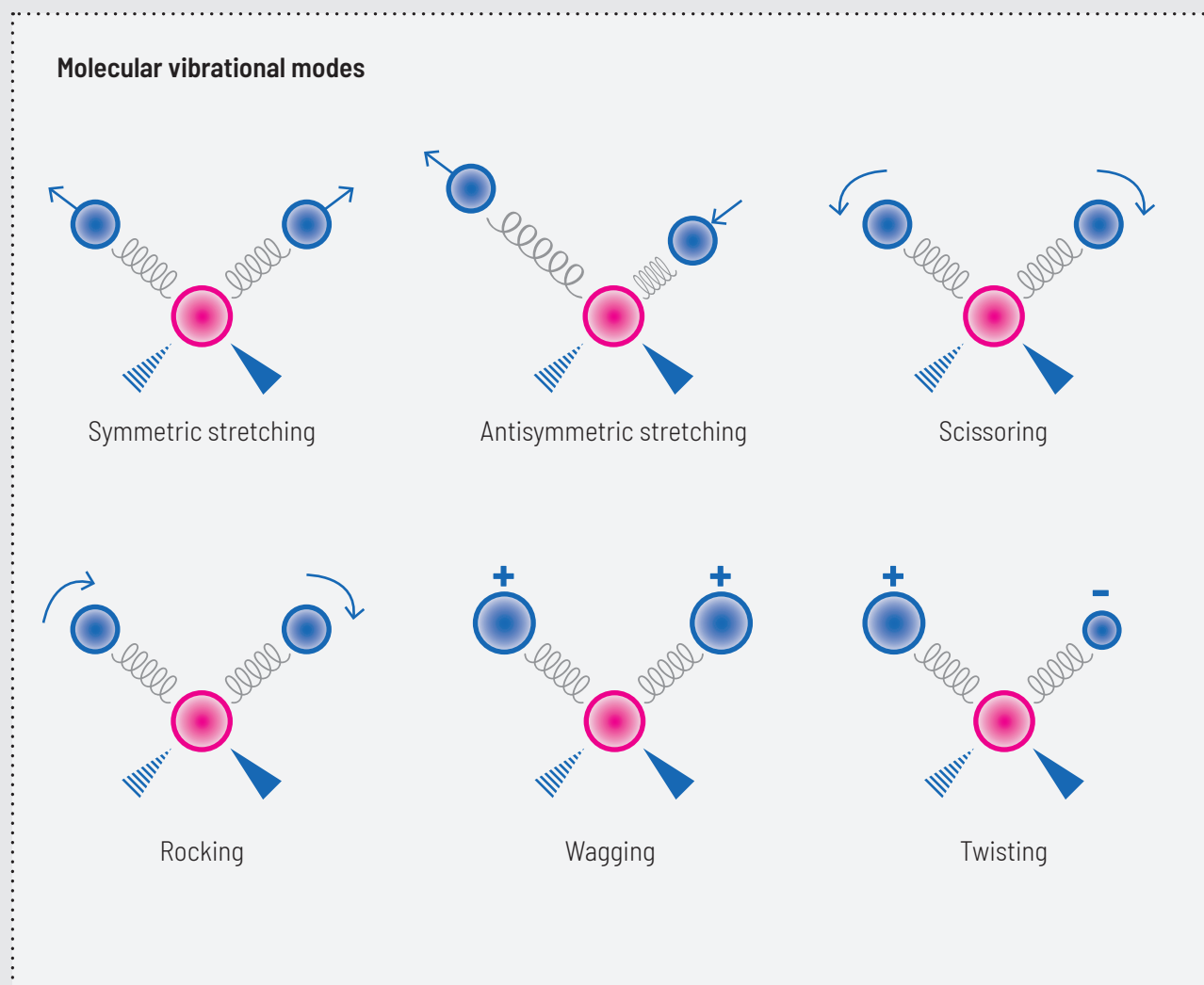
## Infrared Spectroscopy (IR)

Infrared spectroscopy (IR spectroscopy or vibrational spectroscopy) involves the interaction of **infrared radiation**,  $12800 - 10 \text{ cm}^{-1}$  ( $780 \text{ nm} - 1000 \text{ }\mu\text{m}$ ), with **a molecule**.

**The IR region** of the spectrum can be divided into three zones, both from the point of view of applications and devices: **Near IR (NIR)**:  $12800-4000 \text{ cm}^{-1}$ , **Mid IR**:  $4000-400 \text{ cm}^{-1}$ ; **Far IR**:  $400-10 \text{ cm}^{-1}$ , being the Mid-IR region where most of the traditional analytical applications occur.

It covers a range of techniques, mostly based on **absorption** spectroscopy. It relies on the fact that molecules absorb **specific frequencies** that are characteristic of their structure and can be associated with a **particular type of bond**.

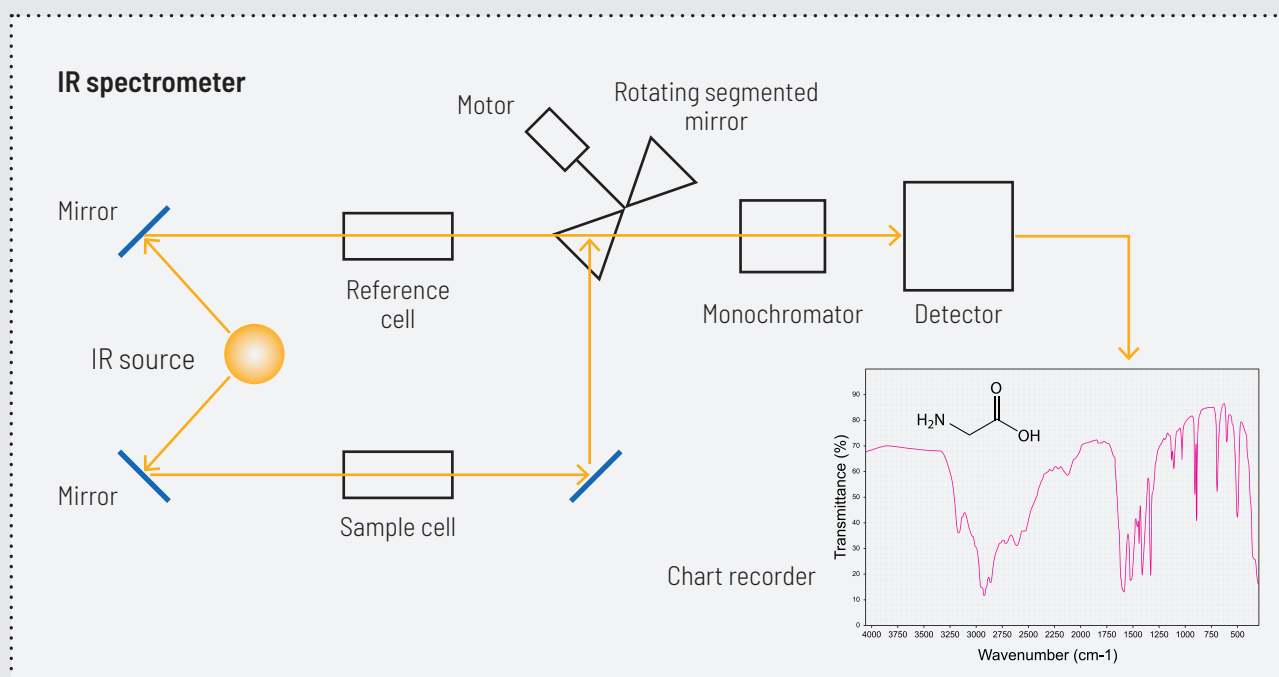
A molecule can **vibrate** in many ways (i.e. symmetric and antisymmetric **stretching**, **scissoring**, **rocking**, **wagging** and **twisting**), and each way is called a vibrational mode and is reflected in the **IR spectrum**. The more complex a molecule is, the more bonds it has and their vibrational spectra are correspondingly more complex.



Infrared (IR) spectroscopy is one of the most common and widely used spectroscopic techniques by inorganic and organic chemists due to its usefulness in **determining structures of compounds** and **identifying them** as each functional group absorbs at a different and specific IR frequency.

## Apparatus

The **spectrophotometers** used to obtain the near and medium infrared consist of a **light source, monochromator** or interferometer and **detector**, which allow obtaining spectra in the region between 780 nm and 25000 nm ( $12800\text{ cm}^{-1}$  to  $400\text{ cm}^{-1}$ ). Currently, infrared spectrophotometers use an interferometer instead of a monochromator in which case the polychromatic radiation impinges on the sample and the spectra are obtained in the frequency domain with the help of the Fourier transform.



Normally the spectrum is given as a function of transmittance, the quotient of the intensity of the transmitted radiation and the incident radiation. It may also be given in absorbance.

The absorbance (A) is defined as the logarithm to base 10 of the reciprocal of the transmittance (T):

$$A = \log_{10} \left( \frac{1}{T} \right) = \log_{10} \left( \frac{I_0}{I} \right)$$

$$T = I/I_0$$

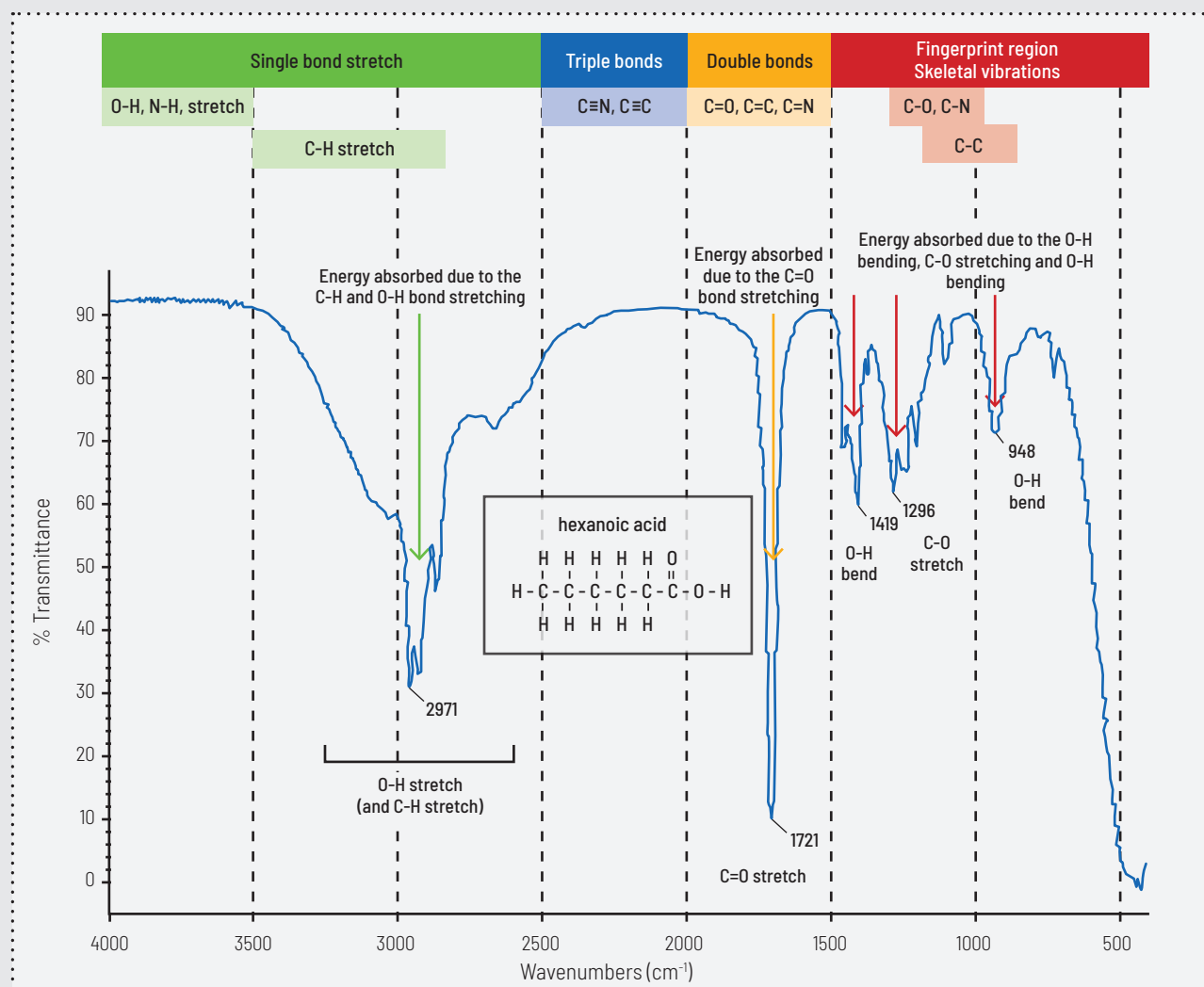
$I_0$  = intensity of incident radiation

$I$  = intensity of transmitted radiation



## IR Spectrum

The following graphic shows an example of an infrared spectrum for hexanoic acid with five different absorptions characteristics of different bond vibrations.



## Sample preparation

IR Spectroscopy is a versatile technique that allows to obtain spectra of solids, liquids and gases using in each case the appropriate cells or supports that must be transparent to the incident radiation. Alkali halides are the most commonly used in the transmission methods (NaCl, KBr, KCl, etc.). In comparison with other instrumental techniques, the samples to be analyzed require little or no preparation.

**Gaseous samples** require a sample cell with a long path length to compensate for the diluteness. The path length of the sample cell depends on the concentration of the compound of interest. A simple glass tube with length of 5 to 10 cm equipped with infrared-transparent windows at the both ends of the tube can be used for concentrations down to several hundred ppm.

**Liquid samples** can be sandwiched between two plates of a salt (commonly sodium chloride, although other salts such as potassium bromide or calcium fluoride are also used). The plates are transparent to the infrared light and do not introduce any lines onto the spectrum.





**Liquid or solid samples in solution** are prepared in a suitable solvent, choosing a concentration and a path length of the cell that give a satisfactory spectrum. Generally, good results are obtained with concentration of 10-100 g/L for a path length of 0.5-0.1 mm. The absorption due to the solvent is compensated by placing in the reference beam a similar cell containing the solvent used.

**Solid samples** are dispersed in a suitable liquid (mull) or in a solid (halide disc), as appropriate. If prescribed in the monograph, make a film of a molten mass between 2 plates transparent to infrared radiation.

## Potassium Bromide for IR Spectroscopy

Potassium bromide (KBr) is the most common alkali halide used as infrared optical windows and components for general spectroscopy because of its transparency from the near ultraviolet to long-wave infrared wavelengths (0.25-25  $\mu\text{m}$ ) and has no significant optical absorption lines in its high transmission region.

In IR spectroscopy, solid samples which are difficult to melt or dissolve in any suitable IR-transmitting solvent are analyzed by grinding with potassium bromide powder, and pressing into a disc.

This method requires a high quality potassium bromide.

In PanReac AppliChem we produce a high quality potassium bromide specially controlled to satisfy the highest requirements in IR spectroscopy.

ANALYSIS CERTIFICATE		
CODE: <b>331489</b>	BATCH:	
PRODUCT: <b>Potassium Bromide for IR</b>		
ISSUE DATE: 07/06/2018	RETEST DATE: 06/2024	
<b>SPECIFICATIONS:</b>	<b>GUARANTEE VALUE</b>	<b>ACTUAL VALUE</b>
Max. Absorbance (2g pressed disc, 0.5 cm layer thickness)	0.004	<0.004
Cl <sup>-</sup> -residue: 2750-3100 cm <sup>-1</sup>	0.05	<0.05
Other: 420-4000 cm <sup>-1</sup>	0.15	<0.15
Water (3200-3700/1620-1640) cm <sup>-1</sup>	passes test	Passes test
Suitability for IR spectroscopy:		



Product name	Code	Package
<b>Potassium Bromide for IR</b>	331489.1608	100 g
	331489.1609	250 g

## Solvents for IR Spectroscopy

The solvent should not exhibit interactions or chemical reactions with the substance to be analyzed, nor should it damage the cell. Those regions of the spectrum in which the solvent has a strong absorption should not be taken into account. The organic solvents to be used must be free of water.

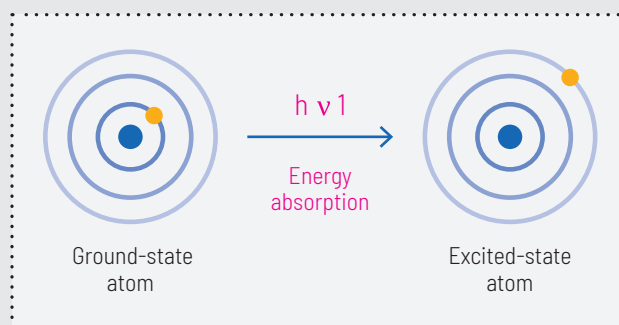
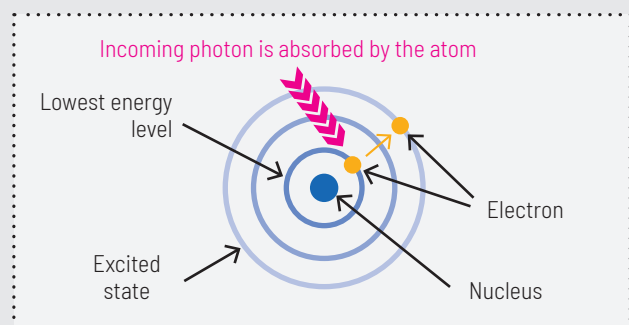
You can find the solvents that PanReac AppliChem can offer in the previous section dedicated to UV spectroscopy.





## Atomic Absorption Spectroscopy (AAS)

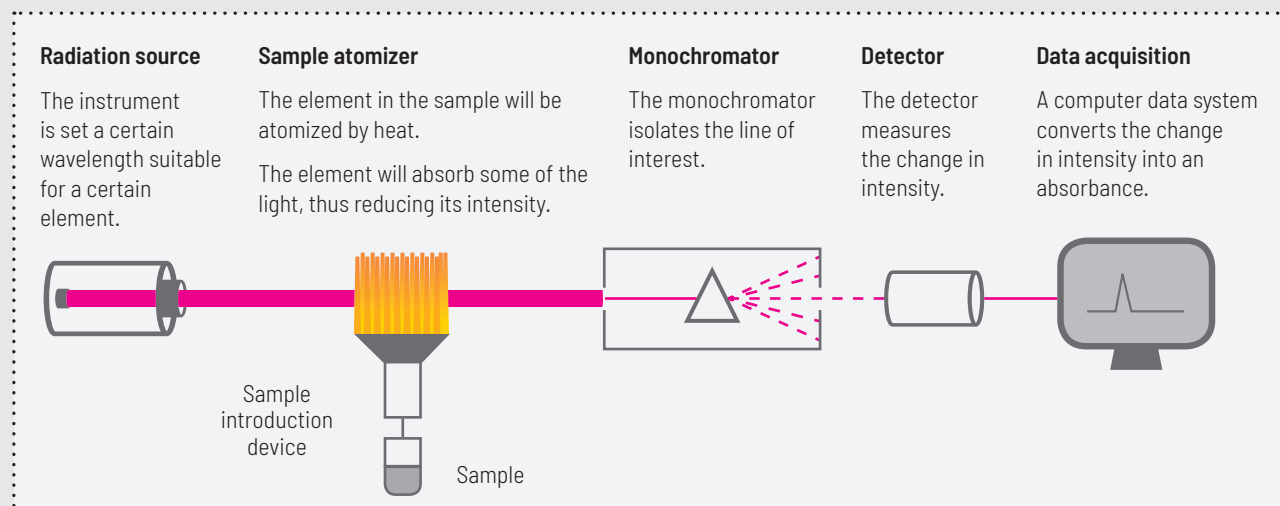
In the **Atomic Absorption Spectroscopy** the visible and ultraviolet light interacts with the atoms. This is a process that occurs when a ground-state atom absorbs electromagnetic radiation of a specific wavelength and is elevated to an excited state. This is a technique for determining the concentration of an element in a sample by measuring the **absorption of electromagnetic radiation** by the **atomic vapor** of the **element** generated from the sample. The amount of radiation absorbed is proportional to the element concentration.



### Apparatus

The apparatus consists essentially of:

- A source of radiation (e.g. hollow-cathode lamps, electrodeless discharge lamps (EDL))
- A sample introduction device
- A sample atomizer
- A monochromator or polychromator
- A detector
- A data acquisition unit



There are 3 types of sample atomizers

- **Flame technique:** a flame atomizer is composed of a **nebulization system** regulation and a **burner**. Fuel-oxidant mixtures are commonly used to produce a range of temperatures from about 2000 K to 3000 K. Samples are nebulized and acidified water is the solvent of choice.
- **Cold vapor and hydride technique:** the atomic vapor may also be generated outside the spectrometer. This is notably the case for cold-vapor method for **mercury** or for certain hydride-forming elements such as **arsenic, antimony, bismuth, selenium** and **tin**. For mercury, atoms are generated by chemical **reduction** with **stannous chloride or sodium borohydride** and the atomic vapor is swept by a stream of an inert gas into a cold quartz cell mounted in the optical path of the instrument. **Hydrides** thus generated are swept by an inert gas into a heated cell in which they are **dissociated into atoms**.
- **Electrothermal atomization technique:** an electrothermal atomizer is generally composed of a **graphite tube furnace** and an **electric power source** atomizes the entire sample and retains the atomic vapor in the light path for an extended period. This **improves the detection limit**. Samples, **liquid as well as solid**, are introduced directly into the graphite tube furnace which is heated in a programmed series of steps to dry the sample and remove major matrix components by pyrolysis and to then atomize all of the analyte.

## Interferences

Chemical, ionization, physical and spectral interferences are encountered in atomic absorption measurements.

- **Chemical interferences** are due to the formation of stable compounds at the combustion temperature of the flame. It can be compensated adding matrix modifiers (as **lanthanum salts**) or releasing agents (as interferences of P, Al or Si).
- **Ionization interferences** are due when the element is easily converted to an ion. Then the number of atoms in ground state able to absorb energy is reduced thus the atomic absorption is reduced. The use of specific ionization buffers (as **Cesium Chloride** or **Potassium Chloride**), which are much easier to ionize, compensates for this effect.
- **Physical interferences** as viscosity or high levels of salts can be eliminated by diluting the sample, for example.
- **Spectral interferences** can be avoided by using a different resonance line.

## Sample preparation

The preparation of the sample may require a dissolution, a digestion (mostly microwave-assisted), an ignition step or a combination thereof in order to clear up the sample matrix and/or remove carbon-containing material.

Sample preparation is always a critical step in atomic spectroscopy analyses. The **reagents used** for sample preparation must be **very pure** in order to make sure to reliably determine the trace elements you are analyzing for.

We offer a **range of acids** of different purity adequate depending on the type of AAS technique used. We also offer **matrix modifiers** to suppress the chemical or ionization interferences, **reduction reagents** (for the cold vapor and hydride techniques) and **standards** traceable to NIST standards reference materials.



# Reagents for Pharma Industry

## Chapter 2



### Acids

Product name	Flame	Cold Vapor	Hydrides	Graphite chamber	Code	Package
<b>Hydrochloric Acid 37%</b>				X	381020.1611	1000 ml
					381020.1612	2.5 L
<b>Hydrochloric Acid 37%</b> (low Hg content)	X	X	X		471020.1611	1000 ml
					471020.1612	2.5 L
<b>Nitric Acid 65%</b>				X	383255.1609	250 ml
					383255.1611	1000 ml
<b>Nitric Acid 65%</b> (low Hg content)	X	X	X		473255.1611	1000 ml
					473255.1612	2.5 L
<b>Sulfuric Acid 95-98%</b> (low Hg content)	X	X	X		471058.1611	1000 ml
					471058.1612	2.5 L
					471058.0716	25 L

### Matrix modifiers

Product name	Chemical suppressor	Ionization suppressor	Code	Package
<b>Buffer Solution Cesium Chloride/Lanthanum Chloride</b>	X	X	176168.1211	1000 ml
<b>Cesium Chloride</b>		X	122509.1206	25 g
			122509.1208	100 g
<b>Lanthanum(III) Chloride 7-hydrate</b>	X		122848.1208	100 g
<b>Lanthanum(III) Nitrate 6-hydrate</b>	X		122669.1208	100 g
<b>Lanthanum(III) Oxide (Reag. Ph. Eur.)</b>	X		122705.1607	50 g
			122705.1609	250 g
<b>Potassium Chloride</b>		X	131494.1210	500 g
			131494.1211	1000 g



### Reducing agents

Product name	Code	Package
<b>Hydroxylammonium Chloride</b>	471914.1209	250 g
<b>Sodium Borohydride for analysis</b>	123314.1606	25 g
	123314.1608	100 g
<b>Tin(II) Chloride 2-hydrate</b>	471303.1609	250 g

### Oxidizing agents

Product name	Code	Package
<b>Potassium Permanganate</b>	471527.1210	500 g

## Standards

Product name	Element (1 g/L)	Composition	Code	Package
Aluminium	Al	$\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ in $\text{HNO}_3$ 0.5N	313170.1208	100 ml
			313170.1210	500 ml
Arsenic	As	$\text{As}_2\text{O}_3$ in $\text{HNO}_3$ 2-5%	313171.1208	100 ml
Barium	Ba	$\text{Ba}(\text{NO}_3)_2$ in $\text{HNO}_3$ 0.5N	313172.1208	100 ml
Cadmium	Cd	$\text{Cd}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$ in $\text{HNO}_3$ 0.5N	313175.1208	100 ml
			313175.1210	500 ml
Calcium	Ca	$\text{Ca}(\text{NO}_3)_2 \cdot 4\text{H}_2\text{O}$ in $\text{HNO}_3$ 0.5N	313176.1208	100 ml
			313176.1210	500 ml
Chromium	Cr	$\text{Cr}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ in $\text{HNO}_3$ 0.5	313179.1208	100 ml
			313179.1210	500 ml
Cobalt	Co	$\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ in $\text{HNO}_3$ 0.5N	313177.1210	500 ml
Copper	Cu	$\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$ in $\text{HNO}_3$ 0.5N	313178.1208	100 ml
			313178.1210	500 ml
Fluoride	F	$\text{NaF}$ in $\text{H}_2\text{O}$	312682.1210	500 ml
Gold	Au	$\text{HAuCl}_4 \cdot 4\text{H}_2\text{O}$ in $\text{HCl}$ 1.2%	313672.1208	100 ml
Iron	Fe	$\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ in $\text{HNO}_3$ 0.5N	313182.1208	100 ml
			313182.1210	500 ml
Lead	Pb	$\text{Pb}(\text{NO}_3)_2$ in $\text{HNO}_3$ 0.5N	313189.1208	100 ml
			313189.1210	500 ml
Lithium	Li	$\text{LiCl}$ in $\text{HCl}$ ~1%	313183.1210	500 ml
Magnesium	Mg	$\text{Mg}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ in $\text{HNO}_3$ 0.5N	313184.1208	100 ml
			313184.1210	500 ml
Manganese	Mn	$\text{Mn}(\text{NO}_3)_2$ in $\text{HNO}_3$ 1N	313185.1208	100 ml
			313185.1210	500 ml
Mercury	Hg	$\text{Hg}(\text{NO}_3)_2$ in $\text{HNO}_3$ 0.5N	313186.1208	100 ml
			313186.1210	500 ml
Molybdenum	Mo	$(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}$ in $\text{NH}_4\text{OH}$ 0.5N	314111.1208	100 ml
			314111.1210	500 ml
Nickel	Ni	$\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ in $\text{HNO}_3$ 0.5N	313187.1208	100 ml
			313187.1210	500 ml
Potassium	K	$\text{KCl}$ in $\text{HCl}$ ~1%	313190.1208	100 ml
			313190.1210	500 ml
Selenium	Se	$\text{SeO}_3\text{H}_2$ in $\text{HNO}_3$ 1N	313191.1208	100 ml
			313191.1210	500 ml
Silicon	Si	$\text{Na}_2\text{SiO}_3$ in $\text{H}_2\text{O}$	312683.1208	100 ml
			312683.1210	500 ml
Silver	Ag	$\text{AgNO}_3$ in $\text{HNO}_3$ 0.5N	313188.1210	500 ml
Sodium	Na	$\text{NaCl}$ in $\text{HCl}$ ~1%	313192.1208	100 ml
			313192.1210	500 ml
Tin	Sn	$\text{SnCl}_4$ in $\text{HCl}$ 5N	313180.1208	100 ml
			313180.1210	500 ml
Titanium	Ti	$\text{TiCl}_4$ in $\text{HCl}$ 5N	313960.1208	100 ml
			313960.1210	500 ml
Zinc	Zn	$\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ in $\text{HNO}_3$ 0.5N	313193.1208	100 ml
			313193.1210	500 ml





## Inductively Coupled Plasma – Atomic Emission Spectrometry (ICP-AES) Mass Spectrometry (ICP-MS)

Inductively Coupled Plasma-Atomic Emission Spectroscopy (**ICP-AES**) also referred to as inductively coupled plasma-optical emission spectrometry (**ICP-OES**), is an **atomic emission spectrometry** method that uses an inductively coupled plasma (ICP) as the excitation source.

Inductively Coupled Plasma-Mass Spectroscopy (**ICP-MS**) is a **mass spectrometry** method that uses an inductively coupled plasma (ICP) as the excitation source.

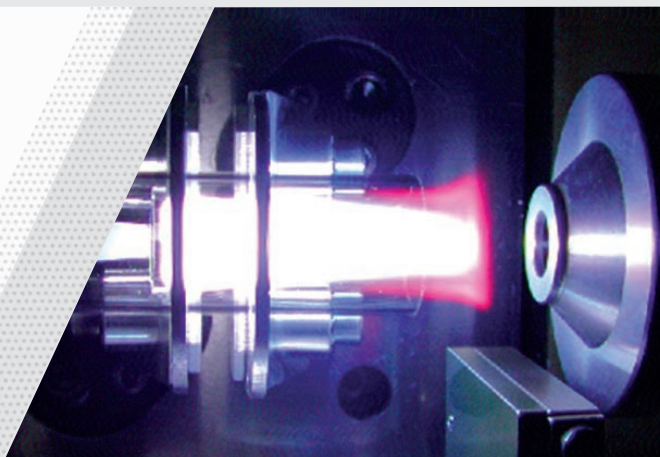


Image courtesy of Perkin Elmer

An ICP is a highly ionized inert gas (usually argon) with equal number of electrons and ions sustained by a radio-frequency (RF) field. The **high temperature** reached in the plasma (7000 °K), successively desolvates, vaporizes, excites (AES detection) and ionizes (MS detection) atoms from the sample.

### ICP-AES

ICP-AES exploits the fact that excited electrons emit energy at a given wavelength as they return to ground state after excitation. The intensity of the **energy emitted** at the chosen wavelength is proportional to the amount (concentration) of that element in the sample being analyzed.

Theoretical detection limit: approximately from 0.1 µg/L (0.1 ppb) to 50 µg/L (50 ppb) depending on the element and assuming there are no spectral interferences.

### ICP-MS

ICP-MS utilizes the ability of the ICP to generate charged ions from the element species within a sample. These ions are then directed into a mass spectrometer, which separates them according to their **mass-to-charge ratio** ( $m/z$ ). Most mass spectrometers have a quadrupole system or a magnetic sector. Ions of the selected mass/charge ratio are directed to a detector (channel electromultiplier, Faraday cup, dynodes), where ion currents are converted into electrical signals.

The element is quantified according to the number of ions arriving and generating electrical pulses per unit time.

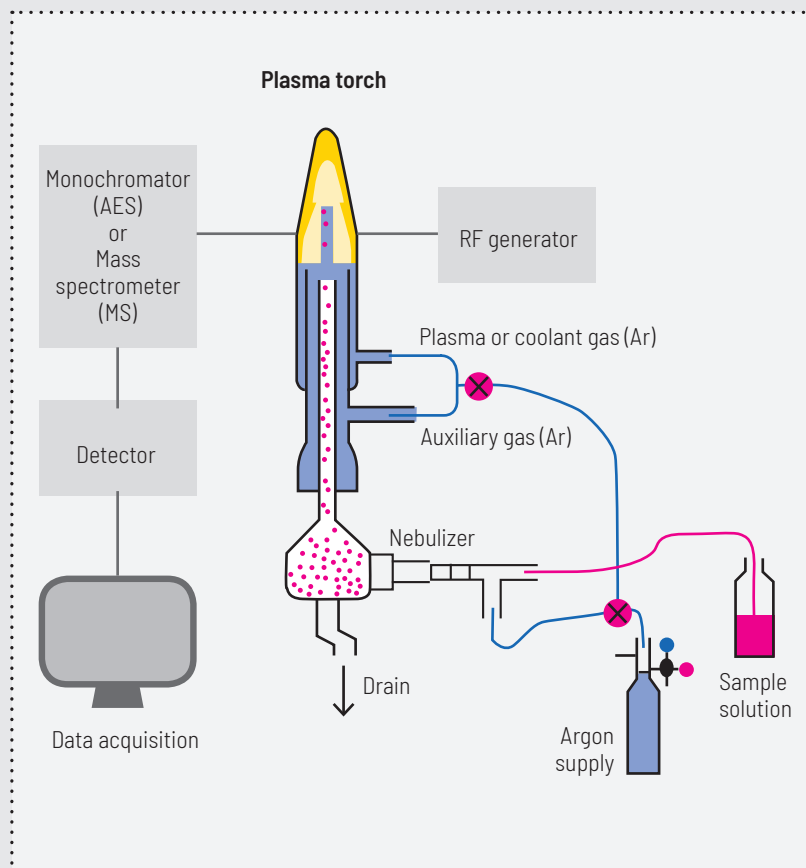
Theoretical detection limit: approximately from 0.001 µg/L (1 ppt) to 0.01 µg/L (10 ppt) depending on the element and assuming there are no spectral interferences.



## Apparatus

The apparatus consists essentially of the following elements:

- Sample introduction system consisting of a peristaltic pump delivering the solution at constant flow rate into nebulizer
- Radio-frequency (RF) generator
- Plasma torch (ICP)
- Transfer optics focusing the image of the plasma at the entrance slit (ICP-AES)
- Wavelength dispersive devices consisting of diffraction gratings, prisms, filters or interferometers (ICP-AES)
- Interface region including cones to transport ions to the ion optics (ICP-MS)
- Mass spectrometer (ICP-MS)
- Detector
- Data acquisition unit



## Sample preparation

The sample preparation usually involves a step of digestion of the matrix by a suitable method, for example in a microwave oven.

Several-sample introduction systems tolerate high acid concentration, but the use of sulfuric ( $\text{H}_2\text{SO}_4$ ) and phosphoric acids ( $\text{H}_3\text{PO}_4$ ) can contribute to background emission. Therefore, nitric ( $\text{HNO}_3$ ) and hydrochloric acids (HCl) are preferable.

The choice of acids used for the preservation of water samples and for the preparation of digested samples is very important. The acid to be used must have a very low concentration of metal in order not to contaminate the samples. Also the choice of the type of acid is very important to avoid:

- unwanted side reactions that can interfere in the analysis,
- formation of volatile compounds which are easily lost from the sample,
- damage to glass parts (special sample introduction system is necessary when using hydrofluoric acid).

Nitric ( $\text{HNO}_3$ ) and hydrochloric (HCl) acids are typically used for the preparation of these samples, often in combination with each other to optimize digestion conditions.

Nitric acid and peroxide ( $\text{H}_2\text{O}_2$ ) are often used for organic matrices as the peroxide is an effective oxidizing agent and destroys the organic matrix.

In some cases hydrofluoric acid (HF) will have to be added to destroy certain materials such as titanium dioxide or silica if they have been used in the final product.

# Reagents for Pharma Industry

## Chapter 2



PanReac AppliChem offers its ranges of **high purity reagents** for trace metal analysis at **ppb and ppt** levels in order to satisfy the high quality demand of reagents for trace analysis by **ICP-OES** in environmental, quality control and research laboratories.

### Acids for trace metal analysis (ppb)

This range consists of acetic, hydrochloric, nitric, sulfuric, hydrofluoric and perchloric acids in 500 mL, 1000 mL and 2.5 L high density polyethylene (HDPE) bottles.

Main advantages:

- In most cases, a metal impurity content of less than 0.1 ppb ( $\mu\text{g/L}$ ) is guaranteed for 60 different elements.

Compared to glass bottles:

- Low level of metal impurities leaching from the polyethylene.
- Light-weight bottles for easier handling and greater safety against breakage in the laboratory.

Product name	Code	Package
Acetic Acid glacial	721008.0011	1000 ml
	721019.0010	500 ml
Hydrochloric Acid 35%	721019.0011	1000 ml
	721019.0012	2.5 L
Hydrofluoric Acid 48%	721028.0010	500 ml
Nitric Acid 69%	721037.0010	500 ml
	721037.0011	1000 ml
	721037.0012	2.5 L
Perchloric Acid 70%	722175.0010	500 ml
	722175.0011	1000 ml
Sulfuric Acid 93-98%	721058.0010	500 ml
	721058.0011	1000 ml



### Acids for trace metal analysis (ppt)

These reagents are **specially purified** by multiple distillation to achieve the **lowest impurity levels**, obtaining products that can be used as a baseline for the analysis of trace metals. The greatest challenge has always been to minimize the impurities of some elements like Fe, Zn, Ni and Cu with a special analytical importance and to limit the presence of elements such as Ca, Na and Al which are ubiquitous in the environment as well as in the container materials. Thanks to this special care in the manufacturing, this range is able to offer reagents with concentrations of the above elements lower than 100 ppt ( $\text{ng/L}$ ) and most lower than 10 ppt ( $\text{ng/L}$ ). It represents the range of reagents with the lowest metal content intended for trace metal analysis. Our reagents for trace metal analysis (ppt) are available in the following sizes in both Teflon PFA and FEP: 250 mL and 500 mL. Water is available in specially washed HDPE bottles.



## Main advantages

- Our reagents for trace metal analysis (ppt) include the analysis of more than 60 metals at parts per trillion levels providing the best level of purity available on the market.
- Manufactured by sub-boiling distillation.
- The concentration level for the majority of metals is below 10 ppt, being all below 100 ppt.
- Batch-to-batch homogeneity for more reproducible results.
- Manufactured and packaged in a clean room, class 10, free of contaminants to guarantee the highest levels of purity.
- Stored in specially selected teflon bottles. The material is controlled prior to the bottle manufacture. Every bottle is leached with hot acid for two weeks in order to eliminate any contamination material due to metal traces.

Product name	Code	Package
<b>Hydrochloric Acid 35%</b>	711019.0009	250 ml
	711019.0010	500 ml
<b>Hydrofluoric Acid 48%</b>	711028.0010	500 ml
<b>Hydrogen Peroxide 30% w/w</b>	716323.0010	500 ml
	711037.0009	250 ml
<b>Nitric Acid 69%</b>	711037.0010	500 ml
	711074.0011	1000 ml
<b>Water</b>	711074.0011	1000 ml



## ICP Standards

For the calibration of the ICP equipment, PanReac AppliChem offers a range of ICP element standards that meet the highest purity and quality requirements demanded by these highly sensitive techniques.

The range of ICP element standards consists of 72 elements in concentrations ranging from 1.000 to 10.000 g/L.

The concentration of each element is strictly controlled to an accuracy of  $\pm 0.2\%$  of the true value, with NIST traceability.

The traces of 69 elements are analyzed by ICP and the result is specified on the certificate of analysis.

The preparation is carried out using raw materials of the highest available purity: 99.999% for metals as well as ultra-pure water and acids. The ICP standards are available in pre-washed 100-mL PE bottles, all of which are supplied with their corresponding detailed certificate of analysis. This certificate includes the composition of the standard, its date of manufacture, the minimum expiry date, the actual value with respect to the guaranteed value, traceability and the concentration of trace impurities.





# Reagents for Pharma Industry

## Chapter 2












### Single Element Standards





Product name	Element	Composition	Code (1 g/L)	Code (10 g/L)	Package
<b>Aluminium</b>	Al	Al in 2-5% HNO <sub>3</sub>	766034.1208	775943.1208	100 ml
<b>Antimony</b>	Sb	Sb in 2-5% HNO <sub>3</sub>	766035.1208	-	100 ml
<b>Arsenic</b>	As	As <sub>2</sub> O <sub>3</sub> in 2-5% HNO <sub>3</sub>	766036.1208	-	100 ml
<b>Barium</b>	Ba	BaCO <sub>3</sub> in 2-5% HNO <sub>3</sub>	766037.1208	-	100 ml
<b>Boron</b>	B	H <sub>3</sub> BO <sub>3</sub> in H <sub>2</sub> O	765900.1208	-	100 ml
<b>Cadmium</b>	Cd	Cd in 2-5% HNO <sub>3</sub>	766038.1208	-	100 ml
<b>Calcium</b>	Ca	CaCO <sub>3</sub> in 2-5% HNO <sub>3</sub>	766040.1208	775950.1208	100 ml
<b>Chloride</b>	Cl	KCl in H <sub>2</sub> O	765903.1208	-	100 ml
<b>Chromium</b>	Cr	Cr in 2-5% HNO <sub>3</sub>	766043.1208	-	100 ml
<b>Cobalt</b>	Co	Co in 2-5% HNO <sub>3</sub>	766041.1208	-	100 ml
<b>Copper</b>	Cu	Cu in 2-5% HNO <sub>3</sub>	766042.1208	-	100 ml
<b>Germanium</b>	Ge	Ge in 2-5% HNO <sub>3</sub> + traces HF	765910.1208	-	100 ml
<b>Gold</b>	Au	Au in 20% HCl	766061.1208	-	100 ml
<b>Iron</b>	Fe	Fe in 2-5% HNO <sub>3</sub>	766049.1208	775967.1208	100 ml
<b>Lead</b>	Pb	Pb in 2-5% HNO <sub>3</sub>	766063.1208	-	100 ml
<b>Magnesium</b>	Mg	Mg in 2-5% HNO <sub>3</sub>	766051.1208	775976.1208	100 ml
<b>Manganese</b>	Mn	Mn in 2-5% HNO <sub>3</sub>	766052.1208	-	100 ml
<b>Mercury</b>	Hg	Hg in 2-5% HNO <sub>3</sub>	766060.1208	-	100 ml
<b>Molybdenum</b>	Mo	Mo in 2-5% HNO <sub>3</sub> + traces HF	766053.1208	-	100 ml
<b>Nickel</b>	Ni	Ni in 2-5% HNO <sub>3</sub>	766054.1208	-	100 ml
<b>Nitrogen</b>	N	(NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub> in H <sub>2</sub> O	765920.1208	-	100 ml
<b>Palladium</b>	Pd	Pd in 20% HCl	765922.1208	-	100 ml
<b>Phosphorus</b>	P	H <sub>3</sub> PO <sub>4</sub> in H <sub>2</sub> O	766384.1208	775907.1208	100 ml
<b>Potassium</b>	K	KNO <sub>3</sub> in 2-5% HNO <sub>3</sub>	766050.1208	775989.1208	100 ml
<b>Rhodium</b>	Rh	RhCl <sub>3</sub> ·3H <sub>2</sub> O in 20% HCl	765926.1208	-	100 ml
<b>Scandium</b>	Sc	Sc <sub>2</sub> O <sub>3</sub> in 2-5% HNO <sub>3</sub>	765930.1208	-	100 ml
<b>Selenium</b>	Se	Se in 2-5% HNO <sub>3</sub>	766055.1208	-	100 ml
<b>Silicon</b>	Si	(NH <sub>4</sub> ) <sub>2</sub> SiF <sub>6</sub> in 5% HF	765997.1208	-	100 ml
<b>Silver</b>	Ag	Ag in 2-5% HNO <sub>3</sub>	766062.1208	-	100 ml

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Product name	Element	Composition	Code (1 g/L)	Code (10 g/L)	Package
<b>Sodium</b>	Na	NaNO <sub>3</sub> in 2-5% HNO <sub>3</sub>	766056.1208	775999.1208	 100 ml
<b>Sulphur</b>	S	H <sub>2</sub> SO <sub>4</sub> in H <sub>2</sub> O	765898.1208	775946.1208	 100 ml
<b>Tin</b>	Sn	Sn in 20% HCl	766047.1208	-	 100 ml
<b>Titanium</b>	Ti	Ti in 5% HF	766057.1208	-	 100 ml
<b>Tungsten</b>	W	W in 5% HF	765939.1208	-	 100 ml
<b>Vanadium</b>	V	V in 2-5% HNO <sub>3</sub>	765938.1208	-	 100 ml
<b>Yttrium</b>	Y	Y <sub>2</sub> O <sub>3</sub> in 2-5% HNO <sub>3</sub>	765915.1208	-	 100 ml
<b>Zinc</b>	Zn	Zn in 2-5% HNO <sub>3</sub>	766058.1208	-	 100 ml
<b>Zirconium</b>	Zr	Zr in 5% HF	765940.1208	-	 100 ml

## Multi Element Standards

Product name	Elements	Matrix	Code	Package
<b>Standard solution containing 9 elements</b>	As, Be, Cd, Cr, Pb, Hg, Ni, Se, Tl (100 mg/L each)	10% HNO <sub>3</sub>	766334.1208	 100 ml
<b>Standard solution containing 21 elements</b>	As, Be, Ca, Cd, Co, Cr, Cu, Fe, Li, Mg, Mn, Mo, Ni, Pb, Sb, Se, Sr, Ti, Tl, V, Zn (100 mg/L each)	5% HNO <sub>3</sub>	766335.1208	 100 ml
<b>Standard solution containing 24 elements</b>	Al, B, Ba, Be, Bi, Ca, Cd, Co, Cr, Cu, Fe, Ga, K, Li, Mg, Mn, Na, Ni, Pb, Se, Sr, Te, Tl, Zn (100 mg/L each)	2% HNO <sub>3</sub>	766333.1208	 100 ml
<b>Standard solution containing 23 elements</b>	Ag, Al, B, Ba, Bi, Ca, Cd, Co, Cr, Cu, Fe, Ga, In, K, Li, Mg, Mn, Na, Ni, Pb, Sr, Tl, Zn (1000 mg/L each)	2% HNO <sub>3</sub>	766332.1208	 100 ml

## Package pictograms

	Glass bottle
	Plastic bottle
	Plastic jerrycan
	Stainless steel drum (returnable)

	Aluminium bottle
	Glass ampoule
	Fluoropolymer bottle



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